NEW STANDARD ACADEMY

SEMRI KOTHI SUPER MARKET, RAEBARELI

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1. Following are the p-block elements:

1.	*Ge*	*As*	*Se*
	32	33	34

If each orbital can take maximum of three electrons and in the in the absence of Aufbau rule, specify the

a)Block b) Group c) Period of the above Elements

2. Elements A,B,C and D have the following electronic configurations:
   a. A: $1s^2,2s^2,2p^1$ B:$1s^2,2s^2,2p^6,3s^2\ 3p^1$
   b. C:$1s^2,2s^2\ 2p^6,3s^2\ 3p^3$ D: $1s^2,2s^2\ 2p^6,3s^2\ 3p^5$
3. Last electron in Lu(71) goes into 5d,but it is studied in *f*-block. Explain
4. Element with electronic configuration $1s^2\ 2s2\ 2p^6\ 3s^2\ 3p^6\ 3d^{10}\ 4s^2\ 4p^6\ 4d^{10}5s^2\ 5p^3$ belongs to which of the following group of the periodic table?
5. In lothar Meyer plot of atomic volume versus atomic mass, the peaks are occupied by
6. A new element discovered has been named Eka-Aluminium.Its atomic number and symbol respectively are
7. Electronic configurations of elements (ground state or excited states) are given
   a) [Ar] $4s^1\ 3d^2$ b) $1s^2\ 2s^2\ sp^5\ 3s^1$
   c) [Ar] $3d^{10}\ 4s^1$ d) [Xe] $4f^{14}\ 6s^2$
8. Without looking at the periodic table select the elements of IIIA group of the periodic table.(atomic numbers are given):
   a) 3,11,19,37 b) 5, 13 21,39
   c) 7,15,31,49 *d) 5, 13, 31,49*
9. Which set does not show correct matching?
   a) Sc3+ [Ne] $3s^2\ 3p^6$ Zero group
   b) Fe2+ [Ar] $3d^6$ VIII group
   c) Cr [Ar] $3d^5 4s^1$ VIBgroup
   d) All of the above
10. Valence electrons in the element A are3 and that in element B are 6.Most probable compound formed formed form A and B is?